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Erratum

IN THE PAPER, "Determination of the Extent of Oxidation of Fats" (J. Am. Oil Chemists' Soc., *34*, 606, 1957), the conversion factor reported for 1 millimol of peroxide to millimols of aldehyde by thermal decomposition of peroxide in rapeseed oil has been found to be too high. By using vacuum and helium instead of pyrogallol-washed nitrogen, which apparently was not oxygen-free, during the thermal decomposition the conversion factor 0.3 was obtained.

The equation given for the calculation of the oxidation degree will be

$$\text{aldehyde value} + 3 \times \text{the peroxide value}$$

The values for the effect of bleaching and of total refining in reducing the content of oxidation products will become 20–40% and 45%, respectively.

• *Letter to the Editor*

SOME TIME AGO I discovered a method for detecting coconut oil in the presence of other fats and oils. I have not seen the reaction mentioned in the literature as a rapid qualitative test for coconut, and I would like to give the details to your Society for comment.

The test consists simply of shaking vigorously a small quantity of the fat or oil with an equal quantity of alcoholic caustic potash. In the presence of coconut oil as little as 5% distinct "fruity" odor results.

The method has been extensively tested in this laboratory, and we have found that in all cases the results have been conclusive.

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ABSTRACTS .... R. A. REINERS, Editor

ABSTRACTORS: Lenore Petschaft Africk, R. R. Allen, S. S. Chang, Sini'tiro Kawamura, F. A. Kummerow, E. G. Perkins, and Dorothy M. Rathmann

• Fats and Oils

BRANCHED-CHAIN FATTY ACIDS. II. ALKALI FUSION OF SOME BRANCHED-CHAIN FATTY ACIDS. R. Lukes and J. Hofman (Czechoslov. akad. věd, Prague). *Collection Czechoslov. Chem. Communs.* 24, 744-54 (1959). (*C. A.* 53, 11202)

INCREASE OF ACTIVITY OF HYDROGENATION CATALYST FOR MAKING HARD ISOÖLEIC ACID. G. I. Kolesnidov (Inst. Food Ind., Krasnodar). *Izvest. Vysshikh Ucheb. Zavedenii, Pishchevaya Tekhnol.* **1958**(5), 48-52. With the same type of catalyst, development of isoölates in hydrogenation of oil is greater the greater the activity of the catalyst. A nickel-copper combination catalyst induced much greater isoölates development than did catalyst derived from nickel formate, even when the catalytic activities were equal. (*C. A.* 53, 10805)

ACTIVITY OF BINARY HYDROGENATION CATALYSTS. B. N. Tyutyunnikov and I. Z. Koshel (Polytech. Inst., Kharky). *Masloboino-Zhirovaya Prom.* 25(2), 14-5 (1959). The productivity of nickel as a catalyst in nickel-cobalt mixture was found to vary with the nickel content of the mixture. At 25% level it was appreciably lower, and at 50% and above slightly higher than the productivity of nickel alone. The low catalytic activity of cobalt was not improved by the addition of copper. (*C. A.* 53, 10805)

NICKEL CATALYST FOR THE HYDROGENATION OF OILS. II. D. D. Nanavati and J. S. Aggarwal (Natl. Chem. Lab., Poona). *J. Proc. Oil Technologists' Assoc. India, Kanpur* 12, 83-91 (1958). Nickel formate catalyst, when supported on neutral carriers of the type, Hyflo-Supercel, decolorizing carbon, or silica gel (*catalyst; carrier* ratio of 2:1) was found to be preferable for increasing the rate of hydrogenation of oils to unsupported catalyst. These neutral supports were superior to acid carriers of the type of fuller's earth or kieselguhr both as regards consistency of hydrogenated product and selectivity of hydrogenation. Supercel was the most satisfactory of the tested supports. The basic carbonate of nickel was also tried and produced satisfactory hydrogenated oils except that they required more time for hydrogenation as compared to the corresponding formate catalyst. (*C. A.* 53, 10805)

SYNTHESIS OF SOME OCTENOIC ACIDS. J. A. Knight and J. H. Diamond (School of Chem. and the Eng. Exper. Station, Georgia Inst. of Techn.). *J. Org. Chem.* 24, 400-03 (1959). The preparation and properties of *trans*-4- and -6- and of *cis*-3-, -4-, and -6-octenoic acids are reported in this paper. The freviously reported isomers, *trans*-3- and *cis*-2-octenoic acids, are also reported in this paper. The *cis* acids were prepared by the catalytic semihydrogenation of the corresponding octynoic acids. The *trans* acids were obtained either directly or indirectly starting with a *trans* alkenoic acid obtained by a Knoevenagel condensation. Physical properties, including infrared spectra, were determined for all of the acids and most of the intermediates. The infrared spectra of the *trans* compounds showed strong absorption in the region of 10.2-10.35 microns. None of the *cis* compounds showed absorption in this region.